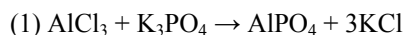
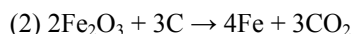


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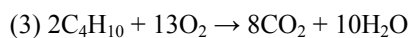
### Mole Ratios



- (a) If 10.2 mol of  $\text{AlCl}_3$  reacts, determine the moles of  $\text{KCl}$  produced
- (b) If 0.40 mol of  $\text{AlCl}_3$  reacts, determine the moles of  $\text{K}_3\text{PO}_4$  reacting
- (c) If 4.3 mol of  $\text{AlPO}_4$  is produced, determine the moles of  $\text{AlCl}_3$  reacting
- (d) If 7.70 mol of  $\text{KCl}$  is produced, determine the moles of  $\text{AlPO}_4$  produced



- (a) If 2.50 mol of  $\text{C}$  reacts, determine the moles of  $\text{Fe}_2\text{O}_3$  reacting
- (b) If 0.04 mol of  $\text{Fe}_2\text{O}_3$  reacts, determine the moles of  $\text{Fe}$  produced
- (c) If 0.300 mol of  $\text{Fe}$  is produced, determine the moles of  $\text{CO}_2$  produced
- (d) If 5.7 mol of  $\text{CO}_2$  is produced, determine the moles of  $\text{C}$  reacting



- (a) If 0.80 mol of  $\text{O}_2$  reacts, determine the moles of  $\text{C}_4\text{H}_{10}$  reacting
- (b) If 0.030 mol of  $\text{C}_4\text{H}_{10}$  reacts, determine the moles of  $\text{CO}_2$  produced
- (c) If 8.3 mol of  $\text{H}_2\text{O}$  is produced, determine the moles of  $\text{O}_2$  reacting
- (d) If 12.4 mol of  $\text{CO}_2$  is produced, determine the moles of  $\text{H}_2\text{O}$  produced

Answers:

(1) (a) 30.6 mol  $\text{KCl}$  produced

(b) 0.40 mol  $\text{K}_3\text{PO}_4$  reacting

(c) 4.3 mol  $\text{AlCl}_3$  reacting

(d) 2.57 mol  $\text{AlPO}_4$  produced

(2) (a) 1.67 mol  $\text{Fe}_2\text{O}_3$  reacting

(b) 0.08 mol  $\text{Fe}$  produced

(c) 0.225 mole  $\text{CO}_2$  produced

(d) 5.7 mol  $\text{C}$  reacting

(3) (a) 0.12 mol  $\text{C}_4\text{H}_{10}$  reacting

(b) 0.12 mol  $\text{CO}_2$  produced

(c) 11 mol  $\text{O}_2$  reacting

(d) 15.5 mol  $\text{H}_2\text{O}$  produced